

1. At constant volume, the mean free path of a gas is

- (a) directly proportional to temperature only
- (b) inversely proportional to pressure
- (c) directly proportional to both temperature and pressure
- (d) independent of temperature and pressure

2. Hydrogen molecule (H_2) with molar mass 2 g mol^{-1} effuses through a small hole at a temperature of 300 K and pressure of 5 torr. Under the same conditions, methane (CH_4) with molar mass 16 g mol^{-1} is also allowed to effuse. The ratio of the rate of effusion of hydrogen (r_{H_2}) and methane (r_{CH_4}) is

- (a) 2.828 : 1
- (b) 1 : 2.828
- (c) 1 : 8
- (d) 8 : 1

3. The work involved in a reversible adiabatic expansion of an ideal gas from p_i and V_i to p_f and V_f is given by

$$1. \quad w = \frac{p_f V_f - p_i V_i}{\gamma - 1}$$

$$2. \quad w = -nC_{v,m} T_i \left[1 - \left(\frac{p_f}{p_i} \right)^{\frac{\gamma}{\gamma - 1}} \right]$$

$$3. \quad w = \frac{p_f V_f - p_i V_i}{1 - \gamma}$$

$$4. \quad w = nC_{p,m} T_i \left[1 - \left(\frac{p_f}{p_i} \right)^{\frac{\gamma}{\gamma - 1}} \right]$$

Select the correct answer using the code given below.

- (a) 1 and 2 only
- (b) 1, 2 and 3 only
- (c) 3 and 4 only
- (d) 1, 2, 3 and 4

4. Consider the following variables :

- 1. Dielectric constant
- 2. Surface tension
- 3. Energy
- 4. Enthalpy

Which one of the following regarding classification of variables is correct?

- (a) 1 and 3 are intensive variables; 2 and 4 are extensive variables
- (b) 1 and 2 are intensive variables; 3 and 4 are extensive variables
- (c) 1 and 4 are intensive variables; 2 and 3 are extensive variables
- (d) 2 and 3 are intensive variables; 1 and 4 are extensive variables

5. Consider the following statements regarding compressibility factor (Z):

Statement-1 :

If a system at state A is converted to state B by two methods (reversible compression and irreversible compression), then the change in internal energy of the system is the same in both the processes.

Statement-2 :

Change in internal energy is a state function.

Which one of the following is correct with respect to the above statements?

- (a) Both Statement-1 and Statement-2 are true and Statement-2 is the correct explanation of Statement-1
 (b) Both Statement-1 and Statement-2 are true and Statement-2 is not the correct explanation of Statement-1
 (c) Statement-1 is true but Statement-2 is false
 (d) Statement-1 is false but Statement-2 is true

6. A 0.04 mole sample of a monatomic gas at 1.00 atm pressure and 298 K expands isothermally and reversibly from an initial volume of 1.0 L to 10.0 L. It is then heated to 500 K, followed by isothermal compression to 1.00 L and finally cooled to 25 °C. The ΔU for the overall process is

- (a) 228 J (b) -228 J
 (c) zero (d) 456 J

7. Consider the following statements :

Statement-1 :

The temperature of a perfect gas is unchanged by Joule-Thomson expansion.

Statement-2 :

For a perfect gas, Joule-Thomson coefficient (μ) is zero.

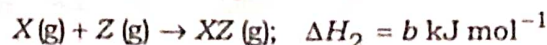
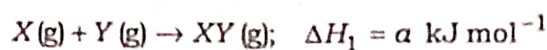
Which one of the following is correct with respect to the above statements?

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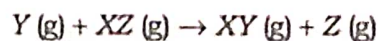
8. In a free expansion process

- (a) the work done is zero and the heat transfer is zero
 (b) the work done is zero but the heat flows into the system
 (c) work is done by the system but the heat transfer is zero
 (d) work is done on the system and the heat flows out of the system

9. Consider the following hypothetical reactions :

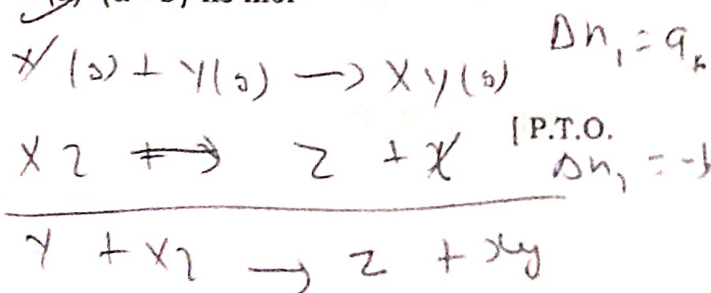


ΔH for the reaction



will be

- (a) $(a+b) \text{ kJ mol}^{-1}$
 (b) $a \text{ kJ mol}^{-1}$
 (c) $b \text{ kJ mol}^{-1}$
 (d) $(a-b) \text{ kJ mol}^{-1}$



$$1 - \frac{300}{400} = 1 - \frac{3}{4}$$

$$\frac{1}{4} \times 100$$

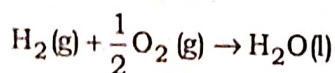
$$\frac{127}{273}$$

$$\frac{273}{400}$$

10. The maximum efficiency of an engine operating between 127 °C and 27 °C is

- (a) 78.7% (b) 25%
(c) 21.25% (d) 75%

11. Consider the following reaction :



The values of ΔH and ΔG are -290 kJ and -230 kJ, respectively, at 300 K. What will be the value of ΔG at 320 K, assuming that there is no variation of ΔH and ΔS with temperature?

- (a) 0.2 kJ
(b) -532 kJ
(c) -226 kJ
(d) -49 kJ

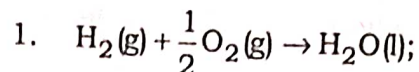
12. Consider the following statements regarding change of enthalpy (ΔH) and change of entropy (ΔS) for a chemical reaction at any temperature T :

1. If ΔH is negative and ΔS is positive, the reaction will be spontaneous at all temperatures.
2. If ΔH is positive and ΔS is negative, the reaction will be spontaneous at low temperatures.
3. When both ΔH and ΔS are positive, the reaction will proceed spontaneously at high temperatures.

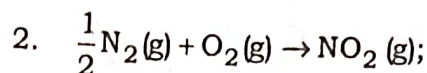
Which of the statements given above is/are correct?

- (a) 1 and 2
(b) 1 and 3
(c) 2 and 3
(d) 2 only

13. Consider the following reactions :



$$\Delta_r H = -285.8 \text{ kJ mol}^{-1}$$



$$\Delta_r H = +33.2 \text{ kJ mol}^{-1}$$

Which one of the following is correct with respect to the above reactions?

- (a) Reaction 1 is endothermic and non-spontaneous at all temperatures
(b) Reaction 2 is exothermic and spontaneous at all temperatures
(c) Reaction 2 is endothermic and non-spontaneous at all temperatures
(d) Reaction 1 is exothermic and spontaneous at all temperatures

14. The experimental results of colligative properties of solutions are often found to differ from the calculated values due to

1. association of the concerned solute
2. dissociation of the concerned solute
3. the solutions being often non-ideal

Select the correct answer using the code given below.

- (a) 1 and 2 only
(b) 2 and 3 only
(c) 1 and 3 only
(d) 1, 2 and 3

15. The vapour pressure of pure water decreases from 24 torr to 23 torr on adding x amount of a solute (molar mass 480 g mol⁻¹) in 150 g of water. The value of x is

- (a) 114 g (b) 128 g
(c) 167 g (d) 1124 g

$$4 \frac{P^\circ - P_s}{P^\circ} =$$

16. The formula for the coordination compound amminediaquachlorido- platinum(IV) tetrachloridoplatinate(II) is

- (a) $[\text{PtCl}(\text{NH}_3)(\text{H}_2\text{O})_2][\text{PtCl}_4]$
 (b) $[\text{PtCl}(\text{NH}_3)(\text{H}_2\text{O})_2]_2[\text{PtCl}_4]_3$
 (c) $[\text{PtCl}(\text{NH}_3)(\text{H}_2\text{O})_2]_3[\text{PtCl}_4]_2$
 (d) $[\text{PtCl}(\text{NH}_3)(\text{H}_2\text{O})_2]_2[\text{PtCl}_4]$

17. How many millilitres of 0.25 M solution of H_2SO_4 will react with 10 mL of 0.25 M solution of NaOH?

- (a) 1 mL
 (b) 2.5 mL
 (c) 5 mL
 (d) 10 mL

$$0.25 \times V = 10 \times 0.25$$

$$V =$$

18. Which one of the following represents the correct order of acidity among CO_2 , SO_2 and SO_3 ?

- (a) $\text{SO}_3 > \text{SO}_2 > \text{CO}_2$
 (b) $\text{SO}_2 > \text{SO}_3 > \text{CO}_2$
 (c) $\text{SO}_2 > \text{CO}_2 > \text{SO}_3$
 (d) $\text{CO}_2 > \text{SO}_2 > \text{SO}_3$

19. Which of the following statements regarding Ostwald's dilution law is/are **not** correct? (Where α = degree of dissociation, k = equilibrium constant for dissociation, c = concentration of solution)

1. Activity coefficients are taken as unity.
 2. Ostwald's law is valid for strong electrolytes.
 3. One may conclude $\alpha = \sqrt{\frac{k}{c}}$.

Select the correct answer using the code given below.

- (a) 1 only
 (b) 2 only
 (c) 1 and 3
 (d) 2 and 3

20. Which one of the following statements is correct?

- (a) Water is strong electrolyte.
 (b) The value of ionic product of water varies with temperature.
 (c) The concentrations of H^+ and OH^- in 10^{-5} M HCl solution are same.
 (d) The ionization constant and ionic product of water are same.

A $6.0 \times 0.2 = 1.2$
 S $4.0 \times 0.2 = 0.8$

4.76
 0.30
 5.06 - 0.48
 4.58

21. The pH of a buffer solution containing 6.0 mL of CH_3COOH (0.2 M) and 4.0 mL of CH_3COONa (0.2 M) solutions will be approximately [where $\text{p}K_a(\text{CH}_3\text{COOH}) = 4.76$; $\log 2 = 0.30$, $\log 3 = 0.48$]

(a) 6.0
 (b) 5.13
 (c) 4.58
 (d) 2.46

$$\text{pH} = \text{p}K_a + \log \frac{[\text{S}]}{[\text{A}]}$$

$$= 4.76 + \log \frac{0.8}{1.2}$$

$$4.76 + \log 2 - \log 3$$

$$4.76 + 0.30 - 0.48$$

22. Which one of the following does **not** represent a correct expression for hydrolysis constant for the equilibrium reaction $\text{A}^- + \text{H}_2\text{O} \rightleftharpoons \text{OH}^- + \text{HA}$?

(a) $K_h = [\text{OH}^-][\text{HA}]/[\text{A}^-]$
 (b) $\text{p}K_h = \text{p}K_w - \text{p}K_a$
 (c) $K_h = K_w / K_a$
 (d) $K_h = [\text{OH}^-][\text{HA}]$

$$K_h = \frac{[\text{OH}^-][\text{HA}]}{[\text{A}^-]}$$

23. Which one among the following compounds does **not** form primary standard solution?

- (a) Na_2CO_3
 (b) NaOH
 (c) ZnO
 (d) Potassium hydrogen phthalate

24. While performing the titration of KMnO_4 with clear Mohr's salt solution, a student obtained brown precipitate in the conical flask instead of a colourless solution, much before the end point was reached. What could be the possible source of error?

- (a) Reaction mixture was not heated to 60°C
 (b) Fe^{2+} in the Mohr's salt solution hydrolyzed to precipitate brown hydroxide
 (c) Reaction medium was not made acidic
 (d) Too large quantity of KMnO_4 had been added to the Mohr's salt solution

25. In which one of the following titrations are metal-ion indicators used?

- (a) Acid-base titration
 (b) Precipitation titration
 (c) Complexometric titration
 (d) Redox titration

26. The total kinetic energy of 2 moles of an ideal gas at 300 K is approximately (where $R = 8.3 \text{ JK}^{-1} \text{ mol}^{-1}$)

- (a) $1.25 \times 10^{-20} \text{ J}$
 (b) $6.3 \times 10^{-21} \text{ J}$
 (c) $7.5 \times 10^3 \text{ J}$
 (d) $1.9 \times 10^3 \text{ J}$

300
 24.9
 12750
 6000
 74700

$\frac{3}{2} \times 2 \times 8.3 \times 300$

27. The root mean square velocity of one mole of a monatomic gas having molar mass M is u_{rms} . The relation between u_{rms} and the average kinetic energy E of the gas is

(a) $u_{rms} = \sqrt{\frac{E}{3M}}$

(b) $u_{rms} = \sqrt{\frac{2E}{M}}$

(c) $u_{rms} = \sqrt{\frac{E}{2M}}$

(d) $u_{rms} = \sqrt{\frac{3E}{2M}}$

$u_{rms} = \sqrt{\frac{3RT}{M}}$
 $\frac{1}{2}mv^2 = E$
 $u = \sqrt{\frac{E}{m}}$

28. Consider the following statements regarding kinetic energy of a gas :

Statement-1 :

At 25 °C, the average kinetic energy of hydrogen gas molecules is the same as that of helium gas atoms.

Statement-2 :

The kinetic energy of a gas depends on the temperature and does not depend on the nature of the gas.

Which one of the following is correct with respect to the above statements?

- (a) Both Statement-1 and Statement-2 are true and Statement-2 is the correct explanation of Statement-1
 (b) Both Statement-1 and Statement-2 are true and Statement-2 is not the correct explanation of Statement-1
 (c) Statement-1 is true but Statement-2 is false
 (d) Statement-1 is false but Statement-2 is true

29. Consider the following van der Waals' constant values of two gases X and Y :

Gas	a kPa dm ⁶ mol ⁻²	b dm ³ mol ⁻¹
X	657	0.04
Y	363	0.11

Which one of the following is correct with respect to the van der Waals' constant values of X and Y in the table given above?

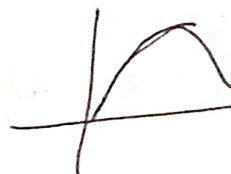
- (a) X is more easily liquefied than Y
 (b) Y is more easily liquefied than X
 (c) Ease of liquefaction does not depend upon the van der Waals' constant a
 (d) Ease of liquefaction does not depend upon the van der Waals' constant b

30. Consider the following statements regarding speed distribution of molecules :

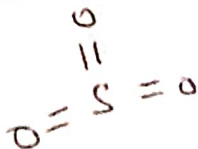
1. The total area under the plots of $\left(\frac{1}{N}\right)\left(\frac{dN_u}{du}\right)$ versus u is a measure of the total number of molecules in the collection.
 2. The curve at any temperature is parabolic near the origin.
 3. The majority of molecules have speeds which cluster around the average value $\bar{u}$ in the distribution curve.

Which of the statements given above is/are correct?

- (a) 1 and 2
 (b) 1 and 3
 (c) 2 and 3
 (d) 3 only



$\frac{3}{2}RT = \frac{3}{2}kT$



31. In SO_3 , one p -orbital and two d -orbitals are used in $\text{S}=\text{O}$ π -bond formation. However, all $\text{S}=\text{O}$ bond lengths are identical. This is due to

(a) involvement of equivalent hybrid orbitals in π -bonding
 (b) considerable polarity in bonds
 (c) delocalization of π -bonds in the molecule
 (d) ionic covalent resonance in the molecule

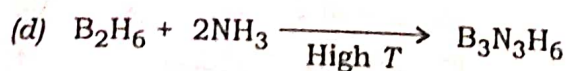
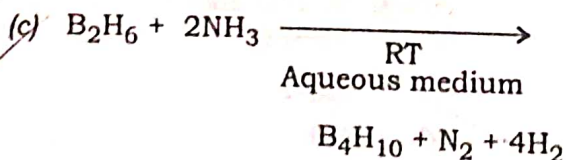
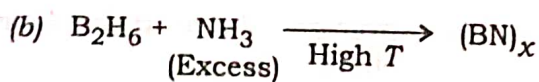
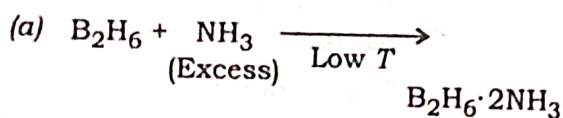
32. Consider the following statements regarding NF_3 :

1. NF_3 has bond angle value much lower than tetrahedral bond angle.
 2. NF_3 is a good Lewis base.

Which of the statements given above is/are correct?

(a) 1 only
 (b) 2 only
 (c) Both 1 and 2
 (d) Neither 1 nor 2

33. Which one of the following reactions is **not** possible between diborane and ammonia?



34. For ZnS , a radius ratio ($r_{\text{Zn}^{2+}}/r_{\text{S}^{2-}}$) of 0.52 suggests a coordination number, 6 for both cation and anion. However, in reality, ZnS adopts a coordination number of 4:4. This is because

(a) ZnS is a true covalent compound

(b) ZnS is an ionic compound with considerable covalency

(c) Zn^{2+} is too small to occupy octahedral holes

(d) S^{2-} ions are too large to provide higher coordination number to Zn^{2+}

35. Theoretically calculated lattice energy (using Born-Landé equation) of hypothetical molecule NaCl_2 is -2180 kJ/mol while that of NaCl is -755 kJ/mol . (Both the molecules are assumed to be formed from elements in their standard states.) Yet NaCl is a stable molecule while NaCl_2 does not exist. The reason for this is

(a) dissociation energy of Cl_2 molecule is too high

(b) second ionization energy of Na metal is too high

(c) electron affinity of Cl atom is too low

(d) Born-Landé equation cannot predict the correct lattice energy of these molecules

36. The solubility of an ionic compound in a solvent depends on

- (a) lattice enthalpy only
- (b) solvation enthalpy only
- (c) ~~lattice enthalpy and solvation enthalpy both~~
- (d) ionization enthalpy of metal

37. Consider the following statements regarding Frenkel defect in solids :

Statement-1 :

Compounds having Frenkel defects have some covalent character.

Statement-2 :

Compounds having Frenkel defects have small positive ions and large negative ions, where small positive ions are highly polarizing and large negative ions are readily polarized.

Which one of the following is correct with respect to the above statements?

- (a) ~~Both Statement-1 and Statement-2 are true and Statement-2 is the correct explanation of Statement-1~~
- (b) Both Statement-1 and Statement-2 are true and Statement-2 is not the correct explanation of Statement-1
- (c) Statement-1 is true but Statement-2 is false
- (d) Statement-1 is false but Statement-2 is true

38. Two types of F atoms in $(\text{CH}_3)_2\text{PF}_3$ are indistinguishable while in $(\text{CH}_3)_3\text{PF}_2$, equatorially placed F atom is distinguishable from apical F atom with the help of ^{19}F NMR spectroscopy. The reason for the same is

- (a) $(\text{CH}_3)_3\text{PF}_2$ has a square pyramidal geometry while $(\text{CH}_3)_2\text{PF}_3$ has a trigonal bipyramidal geometry
- (b) $(\text{CH}_3)_3\text{PF}_2$ has all equivalent hybrid orbitals but $(\text{CH}_3)_2\text{PF}_3$ has non-equivalent hybrid orbitals
- (c) inductive effect of methyl groups is higher in $(\text{CH}_3)_2\text{PF}_3$ than in $(\text{CH}_3)_3\text{PF}_2$
- (d) two methyl groups pose higher rigidity in the molecule to allow pseudorotation in $(\text{CH}_3)_2\text{PF}_3$

39. The bond angle in H_2O molecule is 104.5° . The p -character in non-bonding set of orbitals will be [where $\cos(104.5^\circ) = -0.25$]

- (a) 20%
- (b) 70%
- (c) ~~75%~~
- (d) 80%

40. The wave function

$$\Psi = \Psi_{\text{cov}} + \lambda \Psi_{\text{H}^+ \text{H}^-} + \lambda \Psi_{\text{H}^- \text{H}^+}$$

expresses the phenomenon of

- (a) vibration
- (b) resonance
- (c) conjugation
- (d) orbital

41. Werner confirmed the presence of primary and secondary valencies in the coordination compounds $\text{CoCl}_3 \cdot 6\text{NH}_3$, $\text{CoCl}_3 \cdot 5\text{NH}_3$ and $\text{CoCl}_3 \cdot 4\text{NH}_3$ by

- (a) treating cold aqueous solutions of cobalt complexes with excess of silver nitrate solution
- (b) treating cold aqueous solutions of cobalt complexes with excess of barium nitrate solution
- (c) treating cold aqueous solutions of cobalt complexes with excess of barium chloride solution
- (d) treating cold aqueous solutions of cobalt complexes with excess of sodium nitrate solution

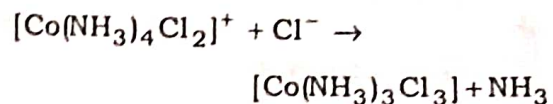
42. Which one among the following chelate rings will be more stable?

- (a) Four-membered chelate ring
- (b) Five-membered chelate ring
- (c) Six-membered chelate ring
- (d) Seven-membered chelate ring

43. Which one of the following is the correct IUPAC name of $\text{Li}[\text{AlH}_4]$?

- (a) Lithium aluminium hydride
- (b) Lithium aluminium tetrahydride
- (c) Lithium tetrahydridoaluminate(III)
- (d) Lithium tetrahydroaluminate(III)

44. Consider the following reaction :



How many isomers of the product complex are possible for the above reaction?

- (a) 1
- (b) 2
- (c) 3
- (d) 5

45. The effective atomic number of $\text{K}_4[\text{Fe}(\text{CN})_6]$ complex is

- (a) 34
- (b) 35
- (c) 36
- (d) 38

46. Which among the following principal quantum numbers has/have five degenerate d -orbitals?

- (a) 2
- (b) 3 only
- (c) 4 only
- (d) Both 3 and 4

47. Which one of the following statements related to radial distribution plots for hydrogen atom is **not** correct?

- (a) Most probable radius varies as $3s > 3p > 3d$.
- (b) Most probable radius varies as $1s < 2s < 3s$.
- (c) Most probable radius varies as $2p < 3p$.
- (d) Most probable radius varies as $3s < 3p < 3d$.

48. Which one of the following sets of quantum numbers represents the highest energy state?

	n	l	m_l	m_s
(a)	4	0	0	$+\frac{1}{2}$
(b)	3	0	0	$-\frac{1}{2}$
(c)	3	1	1	$+\frac{1}{2}$
(d)	3	2	1	$-\frac{1}{2}$

49. Which one of the following groups of elements contains true metals only?

- (a) Carbon, boron, silicon
- (b) Magnesium, aluminium, selenium
- (c) Boron, carbon, nitrogen
- (d) Sodium, thallium, bismuth

50. The electronic configuration of an element A is $[\text{Xe}]4f^1 5d^1 6s^2$ and that of an element B is $[\text{Xe}]4f^3 6s^2$. The elements A and B are

- (a) transition element and lanthanide, respectively

- (b) lanthanide and actinide, respectively
- (c) transition element and actinide, respectively
- (d) both lanthanides

51. The physical and chemical properties of elements are periodic functions of their

- (a) atomic weight
- (b) atomic size
- (c) atomic number
- (d) atomic shape

52. Which one of the following represents the correct order of ionic radii among Y^{3+} , La^{3+} , Eu^{3+} and Lu^{3+} ?

- (a) $\text{Y}^{3+} < \text{La}^{3+} < \text{Eu}^{3+} < \text{Lu}^{3+}$
- (b) $\text{Lu}^{3+} < \text{Eu}^{3+} < \text{La}^{3+} < \text{Y}^{3+}$
- (c) $\text{Eu}^{3+} < \text{La}^{3+} < \text{Lu}^{3+} < \text{Y}^{3+}$
- (d) $\text{Lu}^{3+} < \text{Y}^{3+} < \text{Eu}^{3+} < \text{La}^{3+}$

53. Using Slater's rule, the effective nuclear charge experienced by $3d$ -electron in titanium (atomic number 22) is

- (a) 22.0
- (b) 7.53
- (c) 5.14
- (d) 3.65

54. Atomic size of nitrogen is less than carbon atom, yet carbon shows maximum catenation property while nitrogen does not. The reason for this is

- (a) lone pair-lone pair repulsions in N—N single bond make the bond weaker to allow catenation
- (b) catenation is not dependent on atomic size
- (c) carbon has poor multiple bonds formation ability
- (d) nitrogen exists as an inert gas and therefore does not react

N $0 2s^1 2p^1$
P $3s^2 3p^4$

55. Consider the following statements :

1. Sulfur has a higher ionization energy than phosphorus.
2. Sulfur has a lower ionization energy than phosphorus.
3. Sulfur has a lower electron affinity than chlorine.

Which of the statements given above is/are correct?

- (a) 1 only
(b) 2 only
(c) 1 and 3
(d) 2 and 3

56. Which one of the following represents the correct order of bond angle among H_2O , H_2Te , H_2Se and H_2S ?

- (a) $H_2O > H_2S > H_2Se > H_2Te$
(b) $H_2S > H_2Se > H_2O > H_2Te$
(c) $H_2O > H_2Se > H_2S > H_2Te$
(d) $H_2Te > H_2O > H_2S > H_2Se$

57. Consider the following statements :

1. On moving from left to right in the periodic table, the metallic character of elements decreases.
2. On moving from left to right in the periodic table, the oxides of elements become less basic.

Which of the statements given above is/are correct?

- (a) 1 only
(b) 2 only
(c) Both 1 and 2
(d) Neither 1 nor 2

58. Carbon is able to show allotropic form due to

- (a) its smaller size
(b) its higher electronegativity
(c) its higher ionization enthalpy and unavailability of d -orbitals
(d) catenation and $p\pi-p\pi$ bond formation

59. A neutral atom of an element has two L and five M electrons. What will be the total number of p -electrons in the element?

- (a) 9
(b) 8
(c) 7
(d) 6
- Handwritten orbital diagram:
 $1s^2 2s^2 2p^4 3s^2 3p^6$

60. Consider the following statements :

Statement-1 :

An aqueous solution of NaH_2PO_4 is acidic and that of Na_2HPO_4 is basic in nature.

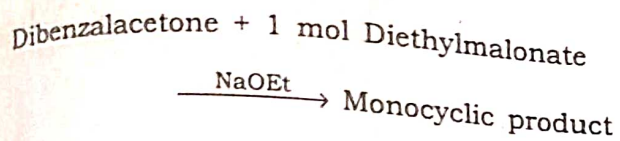
Statement-2 :

An equimolar mixture of NaH_2PO_4 and Na_2HPO_4 acts as a buffer.

Which one of the following is correct with respect to the above statements?

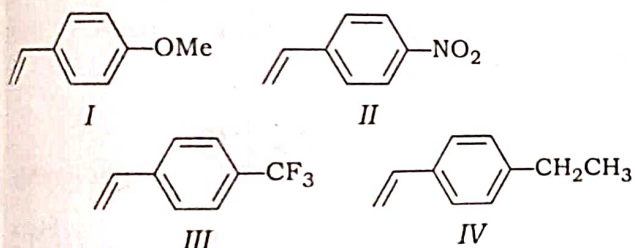
- (a) Both Statement-1 and Statement-2 are true and Statement-1 is not the correct explanation of Statement-2.
(b) Both Statement-1 and Statement-2 are true and Statement-1 is the correct explanation of Statement-2.
(c) Statement-1 is true and Statement-2 is false.
(d) Statement-1 is false and Statement-2 is true.

61. Which one of the following statements is **not** correct regarding the following reaction?



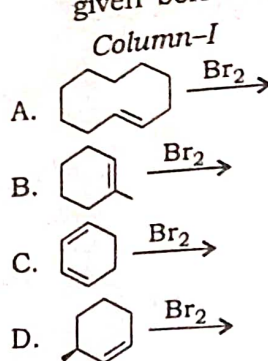
- The product is formed by two successive Michael additions.
- The product is formed by Michael addition followed by intra-aldol condensation.
- Monocyclic product does not give the test of unsaturation with bromine.
- The double bond equivalence of monocyclic product is twelve.

62. Classify the following substituted styrenes with respect to their preferred mode of addition polymerization and select the correct answer accordingly :



- I and IV undergo anionic polymerization while II and III undergo cationic polymerization
- III and IV undergo cationic polymerization while I and II undergo anionic polymerization
- I, III and IV undergo cationic polymerization while II undergoes anionic polymerization
- I and IV undergo cationic polymerization while II and III undergo anionic polymerization

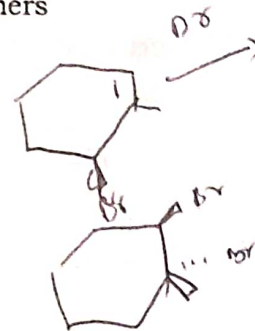
63. Match the reactions from Column-I with the stereochemical outcome of the products described in Column-II and select the correct answer using the code given below the Columns :



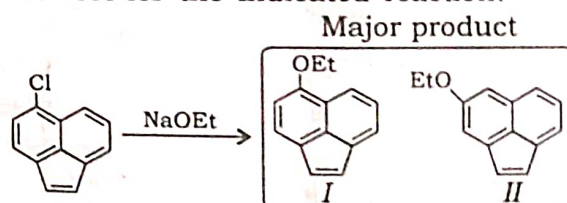
- Column-II
- Formation of diastereomers
 - Formation of racemic mixture
 - Formation of meso-isomer
 - Formation of constitutional isomers

Code :

(a)	A	B	C	D
	2	1	4	3
(b)	A	B	C	D
	2	4	1	3
(c)	A	B	C	D
	3	2	4	1
(d)	A	B	C	D
	4	2	1	2

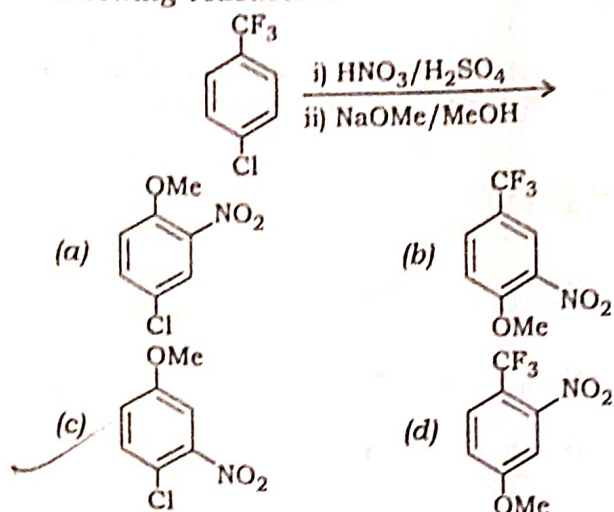


64. Which one of the following statements is correct for the indicated reaction?

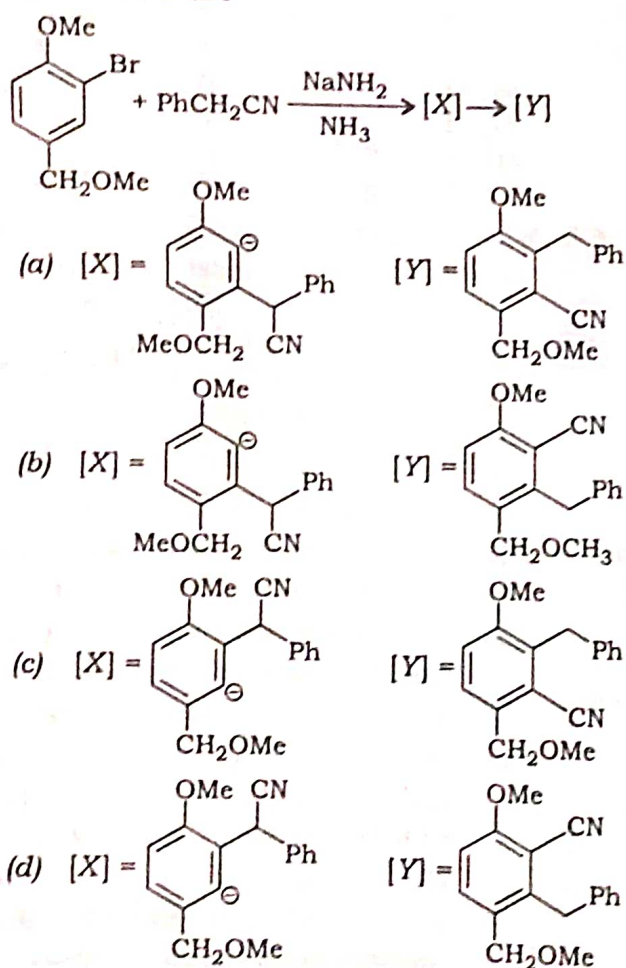


- The major product formed is I and the reaction proceeds through addition-elimination pathway.
- The major product formed is II and the reaction proceeds through addition-elimination pathway.
- The major product formed is I and the reaction proceeds through elimination-addition pathway.
- The major product formed is II and the reaction proceeds through elimination-addition pathway.

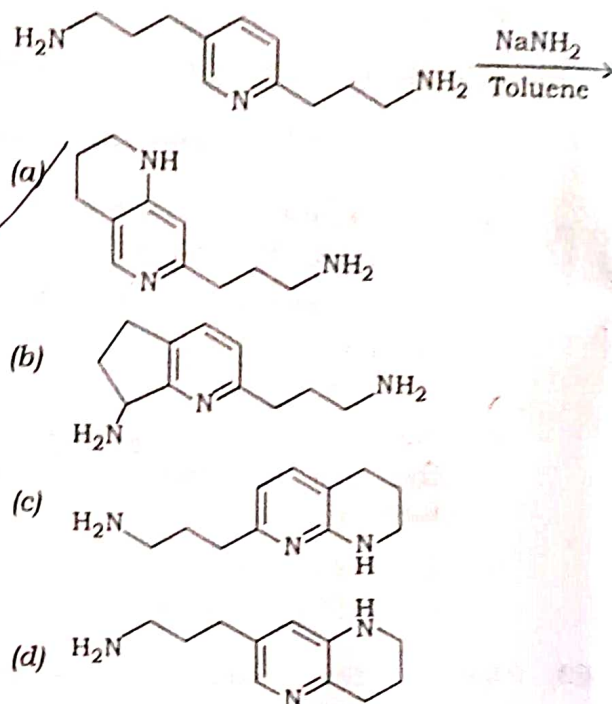
65. The major product formed in the following reaction is



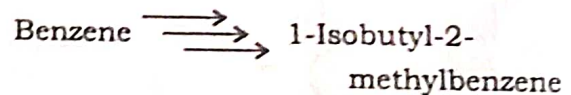
66. The intermediate [X] and the major product [Y] formed in the following reaction are



67. The major product formed in the following reaction is



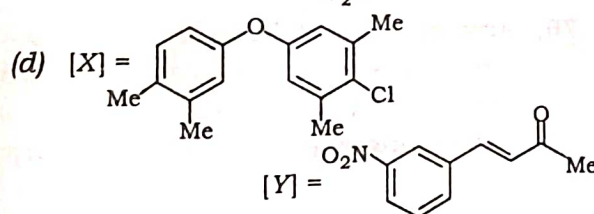
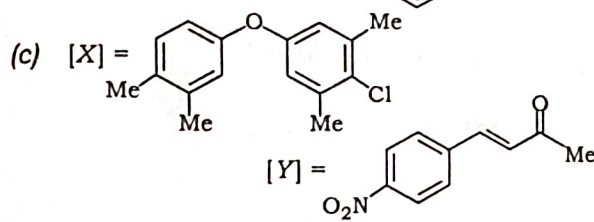
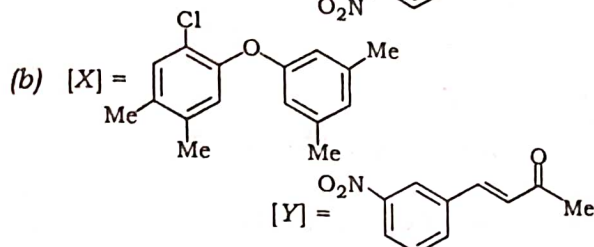
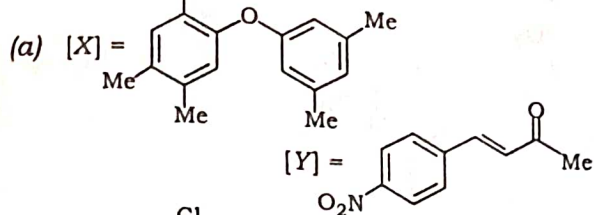
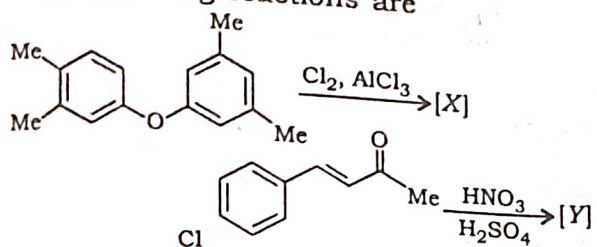
68. The correct sequence of reactions to accomplish the multi-step synthesis of 1-isobutyl-2-methylbenzene from benzene is



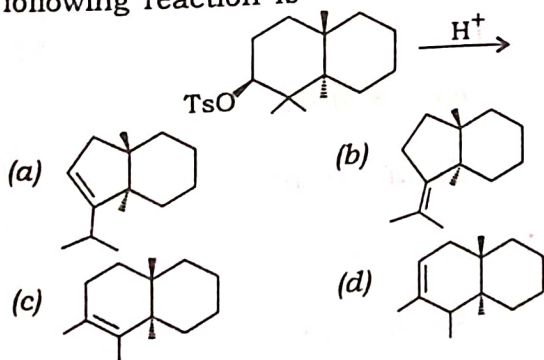
- (a) [1] (CH3)2CHCOCl / AlCl3; [2] HCl, Zn(Hg)/heat; [3] dilute H2SO4; [4] CH3Cl, AlCl3; [5] fuming H2SO4
- (b) [1] CH3Cl, AlCl3; [2] dilute H2SO4; [3] (CH3)2CHCOCl / AlCl3; [4] HCl, Zn(Hg)/heat; [5] fuming H2SO4
- (c) [1] (CH3)2CHCl / AlCl3; [2] fuming H2SO4; [3] CH3Cl, AlCl3; [4] dilute H2SO4
- (d) [1] (CH3)2CHCOCl / AlCl3; [2] HCl, Zn(Hg)/heat; [3] fuming H2SO4; [4] CH3Cl, AlCl3; [5] dilute H2SO4



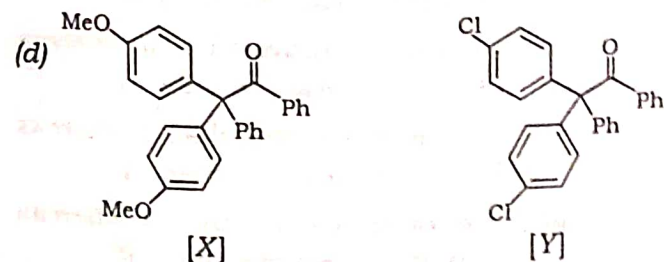
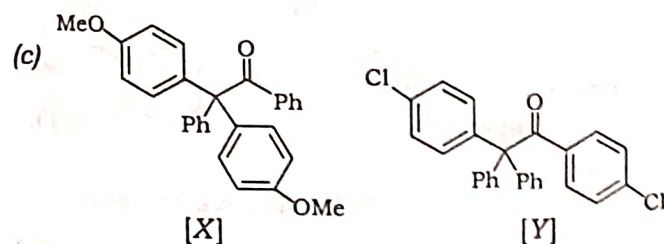
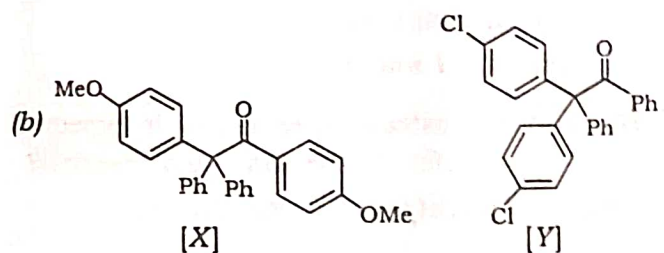
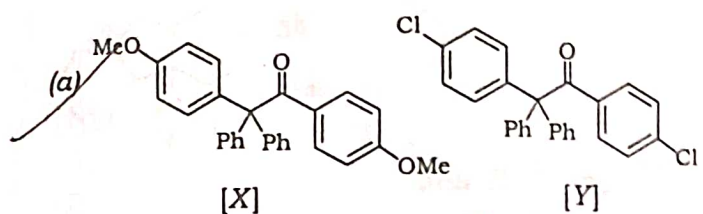
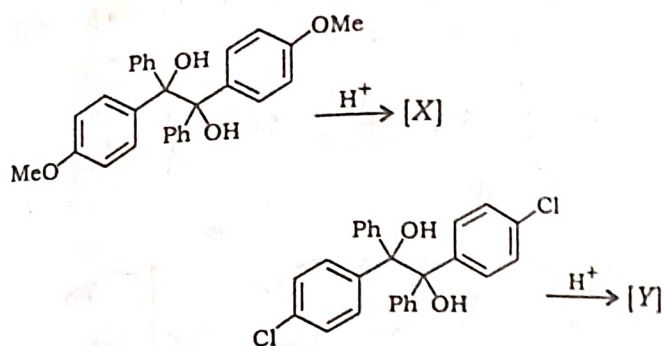
69. The major products [X] and [Y] formed in the following reactions are



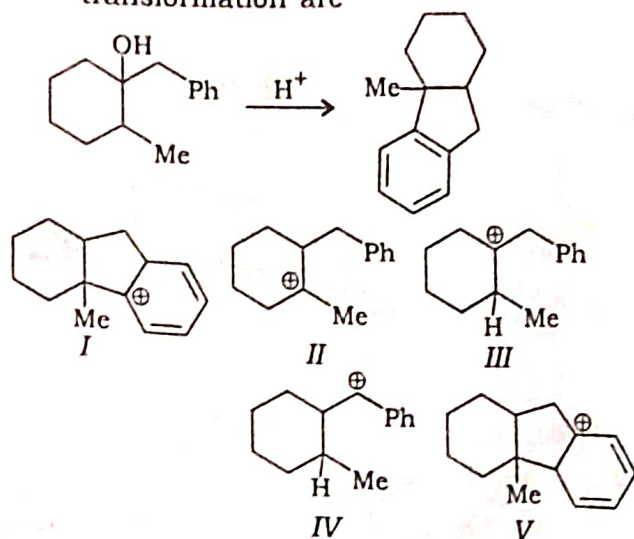
70. The major product formed in the following reaction is



71. The major products [X] and [Y] formed in the following reactions are

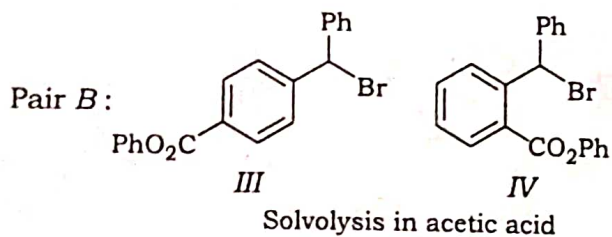
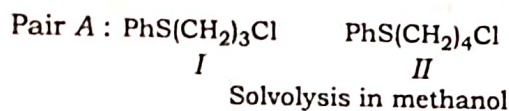


72. The reasonable carbocation intermediates which are likely to form in the reaction mechanism of the following transformation are



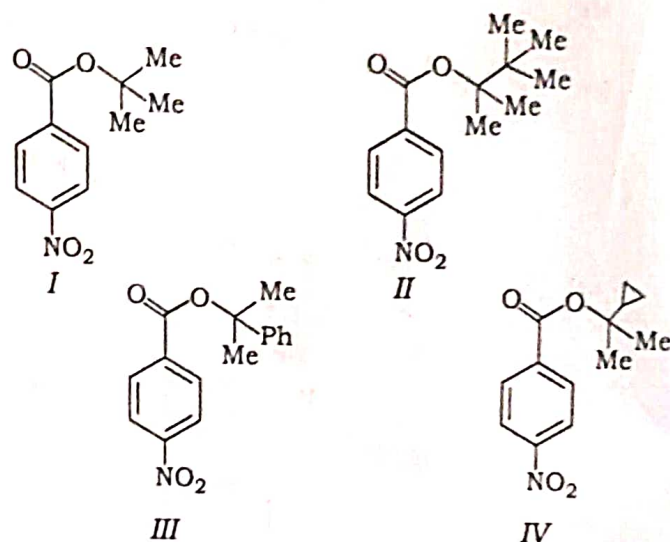
- (a) I, III and IV
(b) III, IV and V
(c) I, II and IV
(d) II, III and V

73. Which substrate in each pair is expected to be faster in the indicated reaction?



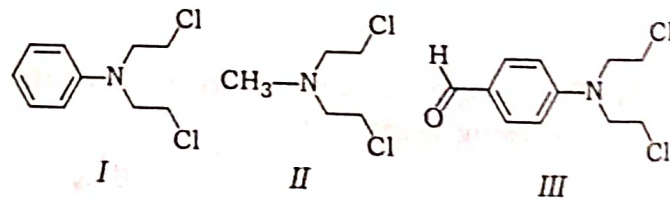
- (a) I is more reactive than II whereas III is more reactive than IV
(b) I is more reactive than II whereas IV is more reactive than III
(c) II is more reactive than I whereas III is more reactive than IV
(d) II is more reactive than I whereas IV is more reactive than III

74. The relative rate of solvolysis in aqueous acetone of the following tertiary *p*-nitrobenzoate esters is



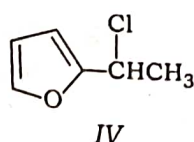
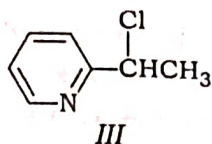
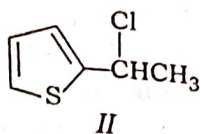
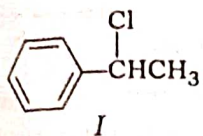
- (a) $I < II < III < IV$
(b) $I < IV < II < III$
(c) $IV < I < III < II$
(d) $II < I < IV < III$

75. Arrange the following three nitrogen mustards in decreasing order of reactivity towards their possible use as alkylating agents for clinical purpose :



- (a) $I > III > II$
(b) $II > I > III$
(c) $III > I > II$
(d) $II > III > I$

76. The increasing order of S_N1 solvolysis reactivity in ethanol of the following substrates is



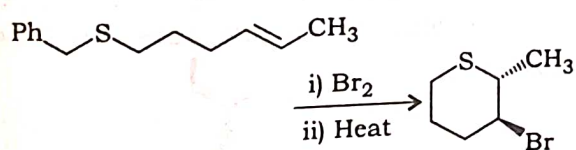
(a) $I < II < IV < III$

(b) $III < I < IV < II$

(c) $III < I < II < IV$

(d) $II < III < IV < I$

77. The number of S_N2 processes involved in the mechanism of the following reaction is



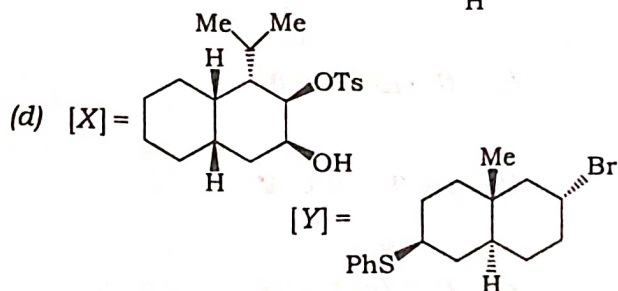
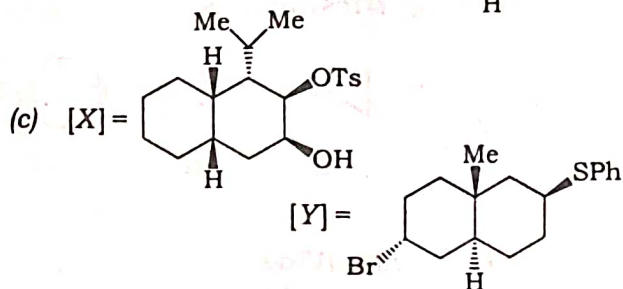
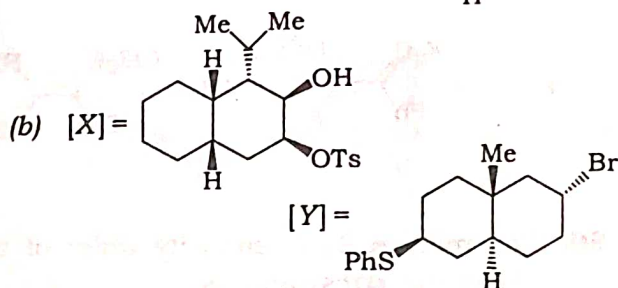
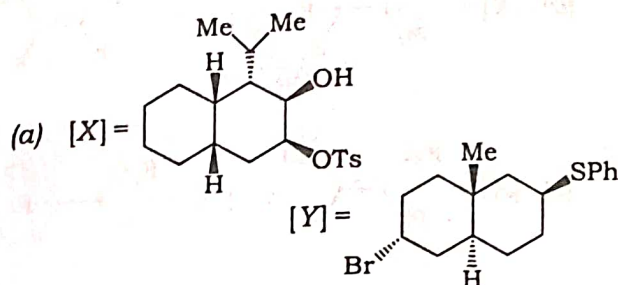
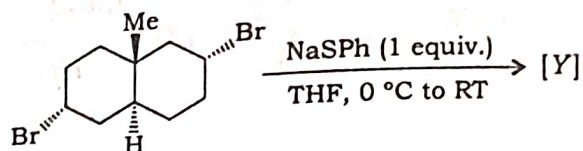
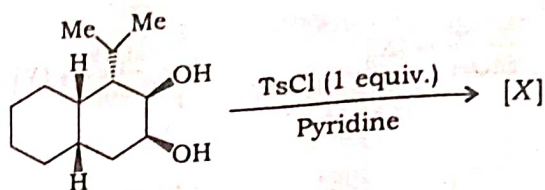
(a) two

(b) three

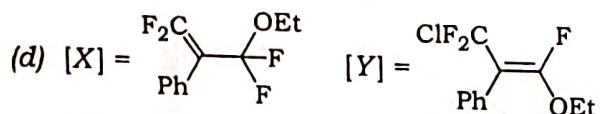
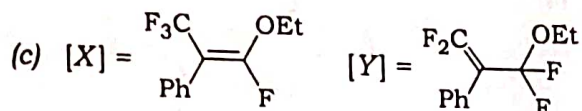
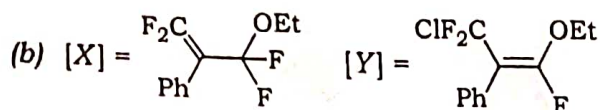
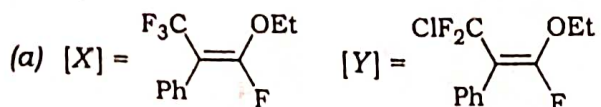
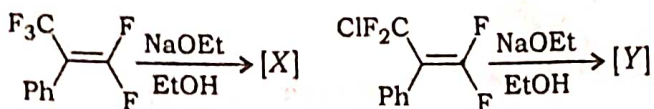
(c) four

(d) five

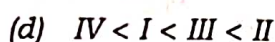
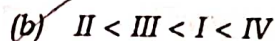
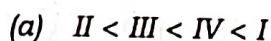
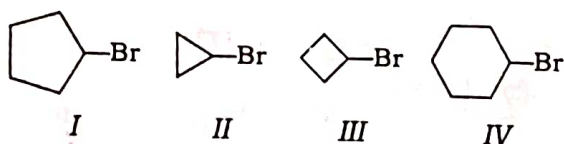
78. The major products [X] and [Y] formed in the following reactions are



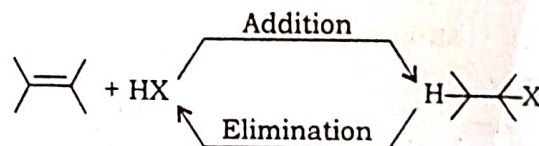
79. The major products [X] and [Y] formed in the following reactions are



80. The relative $\text{S}_{\text{N}}2$ reactivity order of the following substrates is



81. Which of the following statements are correct about the following reactions?

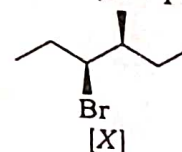


1. Addition is favoured at low temperature whereas elimination is favoured at high temperature.
2. Addition is favoured at high temperature whereas elimination is favoured at low temperature.
3. At high temperature, ΔG is positive for addition reaction whereas ΔG is negative for elimination reaction.
4. At high temperature, ΔG is negative for addition reaction whereas ΔG is positive for elimination reaction.

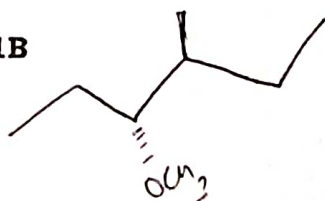
Select the correct answer using the code given below.

- (a) 1 and 3 $\Delta G = \Delta H - T \Delta S$
 (b) 1 and 4
 (c) 2 and 3
 (d) 2 and 4

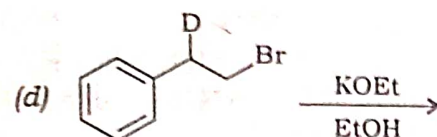
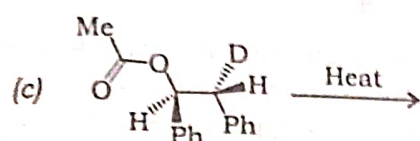
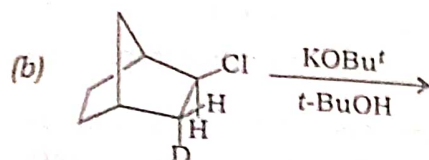
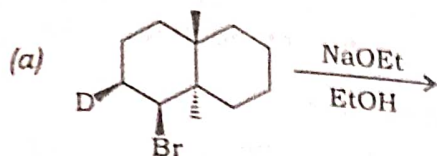
82. The major products formed in $\text{E}2$ and $\text{S}_{\text{N}}2$ reactions of the substrate [X] with sodium methoxide, respectively, are



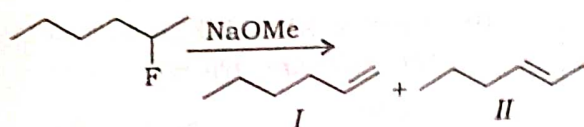
- (Z)-3-methylhex-3-ene and (3R,4R)-3-methoxy-4-methylhexane
- (Z)-3-methylhex-3-ene and (3R,4S)-3-methoxy-4-methylhexane
- (E)-3-methylhex-3-ene and (3R,4S)-3-methoxy-4-methylhexane
- (E)-3-methylhex-3-ene and (3R,4R)-3-methoxy-4-methylhexane



83. In which one of the following elimination reactions, deuterium is **not** retained in the alkene formed as the major product?

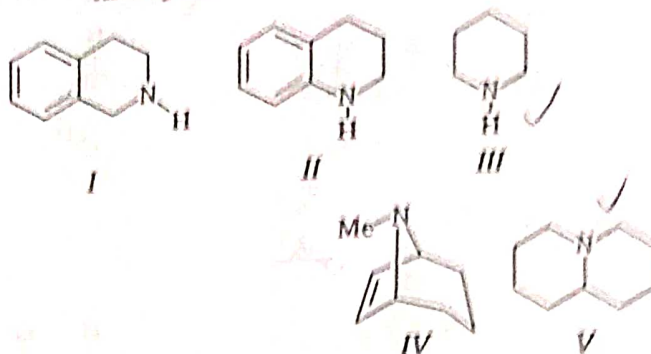


84. Which one of the following statements is correct regarding the given reaction?



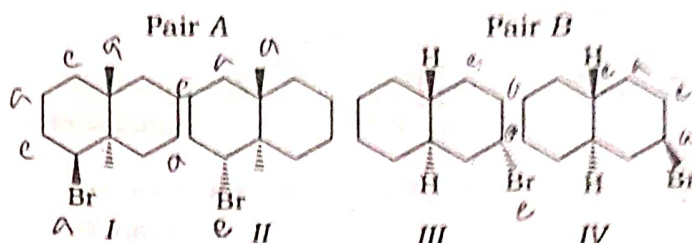
- (a) The major product formed is I and is called Zaitsev product.
 (b) The major product formed is I and is called Hofmann product.
 (c) The major product formed is II and is called Zaitsev product.
 (d) The major product formed is II and is called Hofmann product.

85. Which of the following substrates do **not** form trimethylamine as one of the products in its Hofmann exhaustive methylation method?



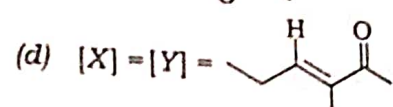
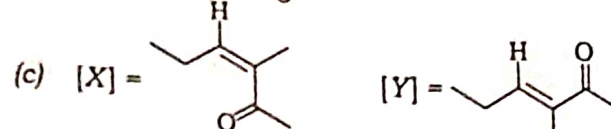
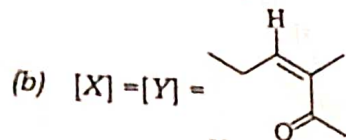
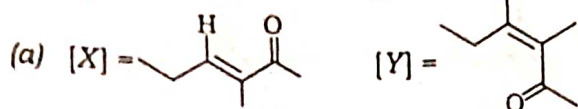
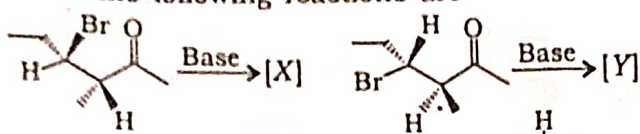
- (a) I, II and III
 (b) I and II only
 (c) III, IV and V
 (d) IV and V only

86. Which of the compounds in Pair A and Pair B are more reactive towards elimination reaction using NaOEt/EtOH?

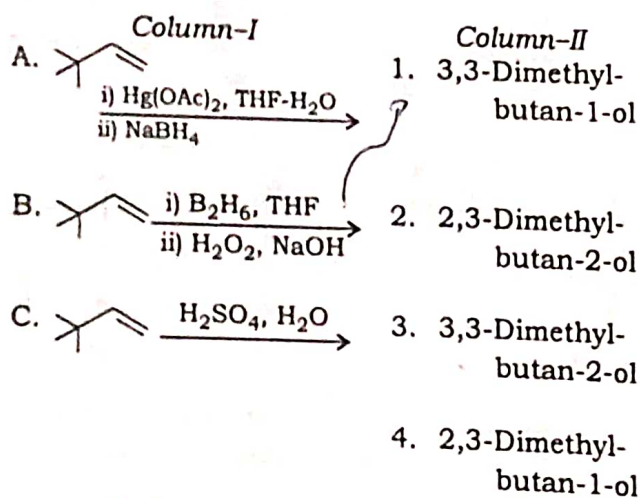


- (a) Pair A : I is more reactive than II
 Pair B : III is more reactive than IV
 (b) Pair A : I is more reactive than II
 Pair B : IV is more reactive than III
 (c) Pair A : II is more reactive than I
 Pair B : III is more reactive than IV
 (d) Pair A : II is more reactive than I
 Pair B : IV is more reactive than III

87. The major products [X] and [Y] formed in the following reactions are



88. Match the reactions from Column-I with the major products formed from Column-II and select the correct answer using the code given below the Columns :



Code :

- | | | | |
|-----|---|---|---|
| (a) | A | B | C |
| | 2 | 4 | 1 |
| (b) | A | B | C |
| | 4 | 3 | 2 |
| (c) | A | B | C |
| | 3 | 1 | 2 |
| (d) | A | B | C |
| | 3 | 4 | 1 |

89. Consider the following statements :

Statement-1 :

The peroxide effect is observed for the addition of HBr to alkene but not for the addition of HCl or HI.

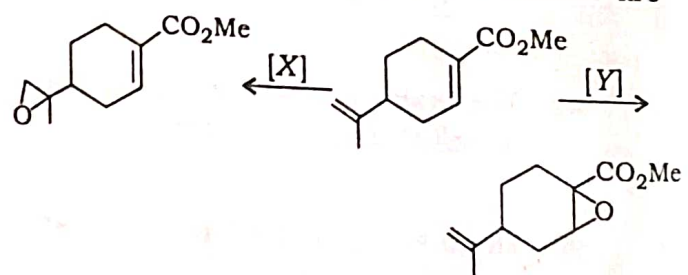
Statement-2 :

Either of the two propagation steps for the addition of HCl or HI to alkene is endothermic in nature.

Which one of the following is correct with respect to the above statements?

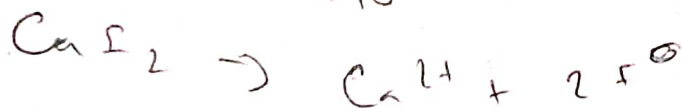
- (a) Both Statement-1 and Statement-2 are correct and Statement-2 is the correct explanation of Statement-1
- (b) Both Statement-1 and Statement-2 are correct and Statement-2 is not the correct explanation of Statement-1
- (c) Statement-1 is correct but Statement-2 is incorrect
- (d) Statement-1 is incorrect but Statement-2 is correct

90. The reagents [X] and [Y] required to carry out the following transformations are



- (a) [X] = *m*-CPBA and [Y] = CH₃CO₃H
- (b) [X] = CH₃CO₃H and [Y] = *m*-CPBA
- (c) [X] = H₂O₂ / NaOH and [Y] = CH₃CO₃H
- (d) [X] = CH₃CO₃H and [Y] = H₂O₂ / NaOH

$$20 \times 10^{-3} = 2 \times 10^{-4}$$



$$(15)(2)^2 = 45^2 \cdot 5 = 45^2$$

91. The conductivity and molar conductance of CaF_2 at 18°C are $4 \times 10^{-5} \text{ S cm}^{-1}$ and $200 \text{ S cm}^2 \text{ mol}^{-1}$, respectively. The K_{sp} for CaF_2 will be

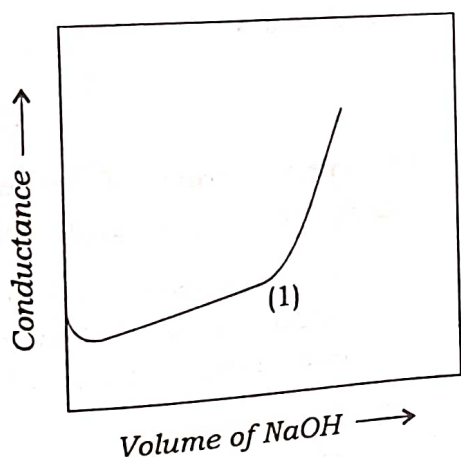
(a) 4.0×10^{-8}

(b) 3.2×10^{-11}

(c) 8.0×10^{-12}

(d) 2.0×10^{-12}

92. In the given plot of conductometric titration of CH_3COOH vs. NaOH , the area around point (1) is curved due to



(a) neutralization of weak acid with weak base

(b) hydrolysis of the salt CH_3COONa

(c) suppression of dissociation of CH_3COOH

(d) poor conductance of CH_3COOH

93. Consider the following table showing the apparent transport numbers of Cd^{2+} ion in CdI_2 solutions of different molalities :

$m/\text{mol kg}^{-1}$	t_+
0.05	0.40
0.1	0.36
1.0	0.30
2.5	0.10
4.0	-0.06
5.0	-0.12
8.0	-0.40
10.0	-0.60

Which one of the following interpretations is correct with respect to the table given above?

(a) In dilute solutions, Cd^{2+} ion migrates towards the anode and I^- migrates towards the cathode and vice versa in concentrated solutions.

(b) In dilute solutions, Cd^{2+} ion migrates towards the cathode and I^- migrates towards the anode and vice versa in concentrated solutions.

(c) In dilute solutions, Cd^{2+} ion migrates towards the cathode and I^- migrates towards the anode and in concentrated solutions, Cd^{2+} ion migrates towards the cathode and CdI_4^{2-} migrates towards the anode.

(d) In dilute solutions, Cd^{2+} ion migrates towards the anode and I^- migrates towards the cathode and in concentrated solutions, Cd^{2+} ion migrates towards the anode and CdI_4^{2-} migrates towards the cathode.

94. The transport number of ions in a 1.0 M solution of AgNO_3 was determined using a 0.3 M solution of $\text{Cd}(\text{NO}_3)_2$ in the moving boundary method. When a current of 15 mA was passed for 30 min, the boundary swept through a volume of 0.1 cm^3 . The transport number of Ag^+ is (Faraday's constant = 96500 C mol^{-1})

(a) 0.36

(b) 0.64

(c) 0.71

(d) 0.83

95. Which of the following electrolytes should be given to a chemist to be used as an indicator electrolyte in the determination of the transport number of Ag^+ in AgNO_3 solution using the moving boundary method?

(a) $\text{Cd}(\text{NO}_3)_2$

(b) LiNO_3 only

(c) CdCl_2 only

(d) Both LiNO_3 and CdCl_2

96. Consider the following statements regarding structures of naphthalene and biphenylene :

1. The C1-C2 bond order is more than C2-C3 bond in naphthalene.
2. The C2-C3 bond order is more than C1-C2 bond in naphthalene.
3. The C1-C2 bond order is more than C2-C3 bond in biphenylene.
4. The C2-C3 bond order is more than C1-C2 bond in biphenylene.

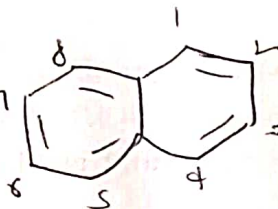
Which of the statements given above are correct?

(a) 1 and 3

(b) 1 and 4

(c) 2 and 3

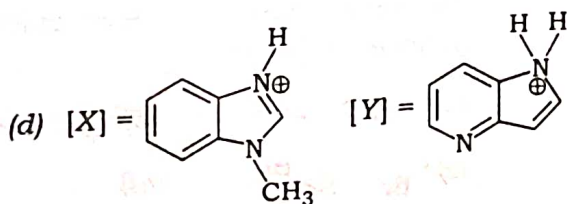
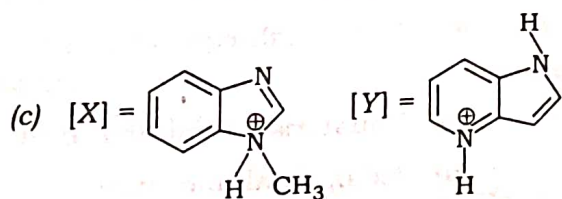
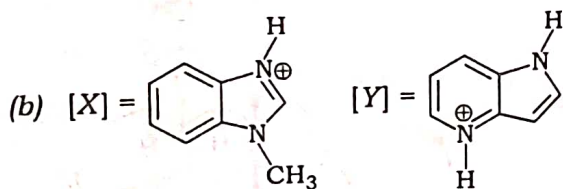
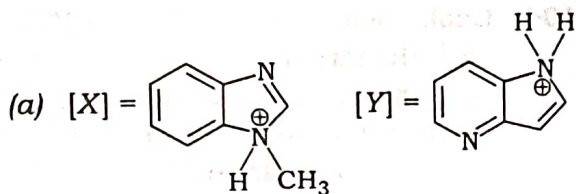
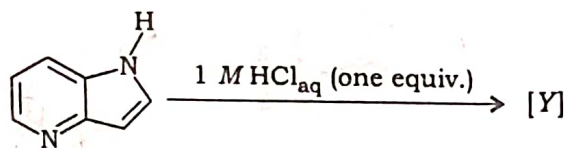
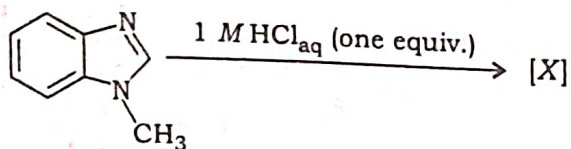
(d) 2 and 4



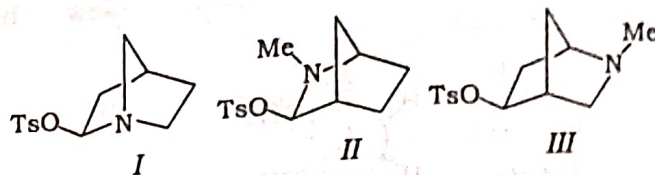
97. The correct order of increasing acid strength of the given carboxylic acids is

- (a) 4-acetylbenzoic acid < 4-cyanobenzoic acid < 4-hydroxybenzoic acid < benzene-1,4-dicarboxylic acid
- (b) 4-hydroxybenzoic acid < 4-acetylbenzoic acid < 4-cyanobenzoic acid < benzene-1,4-dicarboxylic acid
- (c) 4-hydroxybenzoic acid < benzene-1,4-dicarboxylic acid < 4-acetylbenzoic acid < 4-cyanobenzoic acid
- (d) 4-cyanobenzoic acid < 4-acetylbenzoic acid < benzene-1,4-dicarboxylic acid < 4-hydroxybenzoic acid

98. The major products [X] and [Y] formed in the following reactions are



99. The correct decreasing order of reactivity towards ionization under acidic conditions of the following compounds is



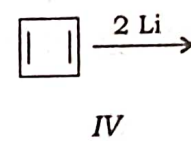
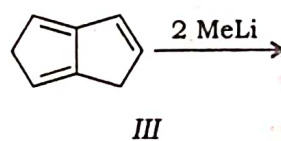
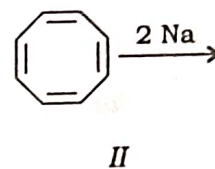
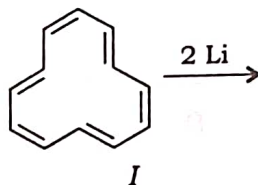
(a) $I > III > II$

(b) $I > II > III$

(c) $III > II > I$

(d) $II > III > I$

100. The products of which of the following reactions are aromatic in nature?



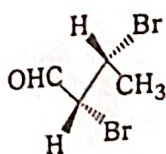
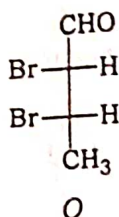
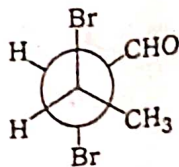
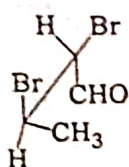
(a) I, II and IV only

(b) II and III only

(c) III and IV only

(d) I, II, III and IV

101. Match the pair of compounds from Column-I with their stereochemical relationship with each other from Column-II and select the correct answer using the code given below the Columns :



Column-I

Column-II

- | | |
|-------------|------------------|
| (A) M and N | 1. Identical |
| (B) M and O | 2. Diastereomers |
| (C) M and P | 3. Enantiomers |
| (D) N and P | |

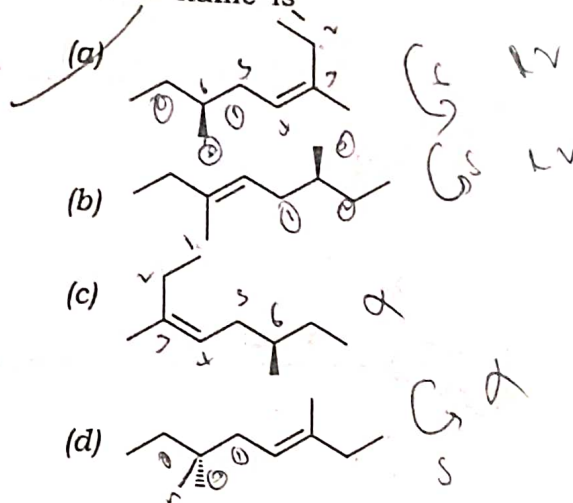
Code :

- | | | | | |
|-----|---|---|---|---|
| (a) | A | B | C | D |
| | 2 | 3 | 2 | 1 |
| (b) | A | B | C | D |
| | 3 | 3 | 1 | 2 |
| (c) | A | B | C | D |
| | 1 | 3 | 1 | 2 |
| (d) | A | B | C | D |
| | 3 | 1 | 2 | 1 |

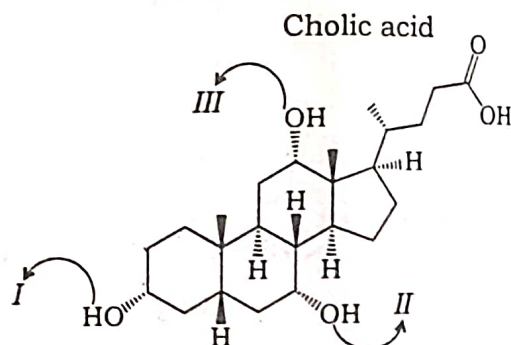
102. Which one of the following has an achiral stereoisomer?

- (a) 2,3-Dibromo-2,3-dimethylbutane
 (b) 2-Bromo-4-methylpentan-3-ol
 (c) 2,3-Dibromopentane
 (d) 1,2-Dibromocyclobutane

103. The structure which is consistent with (R, Z)-3,6-dimethyloct-3-ene as its IUPAC name is

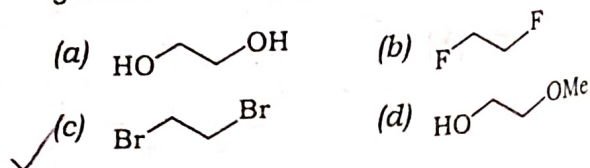


104. Cholic acid, a major constituent of bile, has the structure as shown below. What are the positions of labeled I, II and III hydroxyl groups, respectively, in their chair conformations?



- (a) Equatorial, equatorial and axial
 (b) Equatorial, equatorial and equatorial
 (c) Equatorial, axial and axial
 (d) Axial, axial and axial

105. In which one of the following compounds, the anti-conformer is more stable than gauche conformer?



106. Consider a solution of 50 g of an enzyme dissolved in water to give 2 dm^3 of solution at 27°C . The osmotic pressure of the given solution is (where molar mass of enzyme = 90000 g mol^{-1} and $R = 0.083 \text{ dm}^3 \text{ bar K}^{-1} \text{ mol}^{-1}$)

- (a) 0.49 kPa
- (b) 0.69 kPa
- (c) 0.89 kPa
- (d) 1.09 kPa

107. Consider the following values of normal boiling point and molal boiling point elevation constant for some solvents :

Solvent	Normal boiling point (K)	Molal boiling point elevation constant $\left(\frac{K_b}{\text{K kg mol}^{-1}} \right)$
Water	373.15	0.52
Ethanol	351.5	1.20

Which one of the following solutions would result in the greatest boiling point elevation?

- (a) 0.05 M lithium chloride in water
- (b) 0.05 M lithium chloride in ethanol
- (c) 0.05 M magnesium chloride in water
- (d) 0.05 M magnesium chloride in ethanol

108. When a non-volatile solute is dissolved in a volatile solvent, the depression in the freezing point is

- (a) inversely proportional to the molality of the solution
- (b) directly proportional to the molality of the solution
- (c) independent of the amount of solute present in the solution
- (d) directly proportional to the mole fraction of the solvent present

109. Consider the following statements regarding Henry's law :

1. The Henry's law constant depends on the nature of the gas.
2. CO_2 is more soluble than O_2 .
3. The Henry's law constant does not depend on the nature of the solvent.
4. The Henry's law constant depends on the temperature.

(The Henry's law constant values of CO_2 and O_2 in water at 298 K are $3.01 \times 10^3 \text{ kPa kg mol}^{-1}$ and $7.92 \times 10^4 \text{ kPa kg mol}^{-1}$, respectively)

Which of the statements given above are correct?

- (a) 1 and 2 only
- (b) 1, 2 and 3
- (c) 1, 2 and 4
- (d) 3 and 4

110. The value of the Henry's law constant of O_2 in water at 298 K is $8.0 \times 10^4 \text{ kPa kg mol}^{-1}$. When its partial pressure is 20 kPa, the molar concentration of O_2 in water at 25 °C is

- (a) 25 mmol dm^{-3}
- (b) $0.25 \text{ mmol dm}^{-3}$
- (c) $0.4 \times 10^{-3} \text{ kg mol}^{-1}$
- (d) $4 \times 10^{-3} \text{ mol kg}^{-1}$

111. A conductivity cell was calibrated using 0.1 M KCl solution and the measured resistance was 100 Ω . The resistance of the same cell when filled with 0.02 M KCl solution is 500 Ω . The molar conductivity of 0.02 M KCl solution will be (the conductivity of 0.1 M KCl solution is 1.3 S m^{-1})

- (a) $13.0 \text{ S cm}^2 \text{ mol}^{-1}$
- (b) $130.0 \times 10^{-4} \text{ S cm}^2 \text{ mol}^{-1}$
- (c) $130.0 \text{ S cm}^2 \text{ mol}^{-1}$
- (d) $260.0 \text{ S cm}^2 \text{ mol}^{-1}$

112. The molar conductivities at infinite dilution Λ_m^∞ for HNO_3 , CH_3COOH and $NaNO_3$ are $421.2 \text{ S cm}^2 \text{ mol}^{-1}$, $391.0 \text{ S cm}^2 \text{ mol}^{-1}$ and $121.5 \text{ S cm}^2 \text{ mol}^{-1}$, respectively. The Λ_m^∞ for CH_3COONa will be

- (a) $51.5 \text{ S cm}^2 \text{ mol}^{-1}$
- (b) $91.3 \text{ S cm}^2 \text{ mol}^{-1}$
- (c) $151.5 \text{ S cm}^2 \text{ mol}^{-1}$
- (d) $201.0 \text{ S cm}^2 \text{ mol}^{-1}$

113. The correct order of molar conductance at infinite dilution of LiCl, NaCl and KCl is

- (a) $LiCl > NaCl > KCl$
- (b) $KCl > NaCl > LiCl$
- (c) $NaCl > KCl > LiCl$
- (d) $KCl > LiCl > NaCl$

114. Consider the following statements regarding ionic molar conductivities :

Statement-1 :

With increasing dilution, the conductivity keeps on increasing.

Statement-2 :

With increasing dilution, the degree of ionization of weak electrolytes increases, and the mobility of ions also increases.

Which one of the following is correct with respect to the above statements?

- (a) Both Statement-1 and Statement-2 are true and Statement-2 is the correct explanation of Statement-1
- (b) Both Statement-1 and Statement-2 are true and Statement-2 is not the correct explanation of Statement-1
- (c) Statement-1 is true but Statement-2 is false
- (d) Statement-1 is false but Statement-2 is true

115. Consider the following statements regarding ionic molar conductivities :

Statement-1 :

The Walden's rule holds good for tetramethylammonium ion and picrate ion.

Statement-2 :

The effective radii of tetramethylammonium ion and picrate ion in the solvent are larger and very different from the crystallographic radii of the ions.

Which one of the following is correct with respect to the above statements?

- (a) Both Statement-1 and Statement-2 are true and Statement-2 is the correct explanation of Statement-1
- (b) Both Statement-1 and Statement-2 are true and Statement-2 is not the correct explanation of Statement-1
- (c) Statement-1 is true but Statement-2 is false
- (d) Statement-1 is false but Statement-2 is true

116. Consider the following statements regarding movement of ions :

- 1. Ionic mobility is inversely proportional to the potential gradient.
- 2. Ionic mobility is directly proportional to the ionic speed.

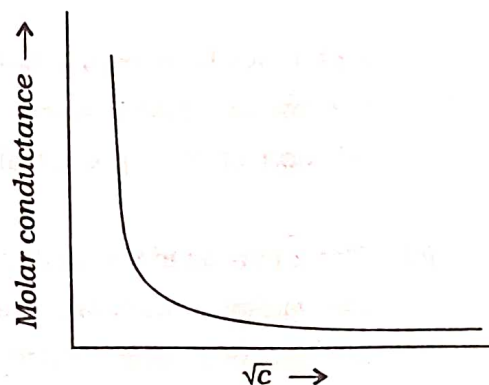
3. Transport number of the ions is inversely proportional to the ionic speed.

4. Sum of the transport numbers of ions in a solution is equal to one.

Which of the statements given above are correct?

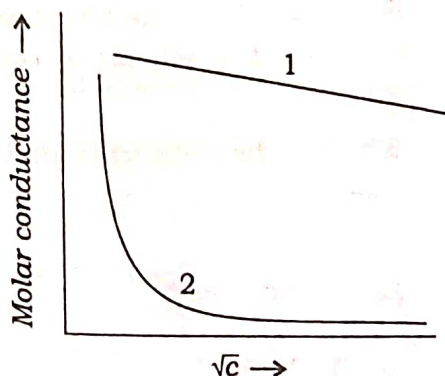
- (a) 1 and 2 only
- (b) 1, 2 and 3
- (c) 3 and 4
- ☒ (d) 1, 2 and 4

117. The following graph shows the variation of molar conductance with the concentration of a/an



- (a) strong acid
- ☒ (b) weak acid
- (c) buffer
- (d) ethanol

118. Consider the following graph showing the variation of molar conductance with concentration :



Which one of the following statements is correct with respect to the above graph?

- (a) Plot 1 can be extrapolated to obtain the molar conductance at infinite dilution of weak electrolytes.
- (b) Plot 1 can be extrapolated to obtain the molar conductance at infinite dilution of strong electrolytes.
- (c) Plot 2 can be extrapolated to obtain the molar conductance at infinite dilution of strong electrolytes.
- (d) Both plots 1 and 2 can be extrapolated to obtain the molar conductance at infinite dilution of strong and weak electrolytes, respectively.

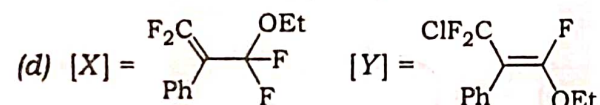
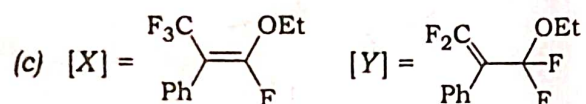
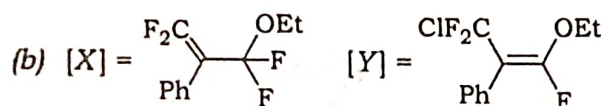
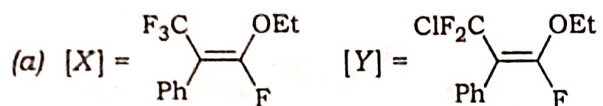
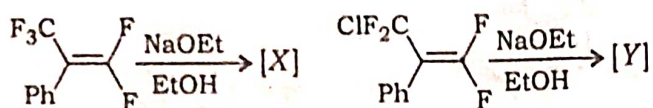
119. The term $A\sqrt{c}$ in the Onsager equation $\Lambda = \Lambda^\infty - (A + B\Lambda^\infty)\sqrt{c}$ represents the

- (a) decrease in molar conductivity due to the electrophoretic effect
- ~~(b) decrease in molar conductivity due to the asymmetry effect~~
- (c) increase in molar conductivity due to the electrophoretic effect
- (d) increase in molar conductivity due to the asymmetry effect

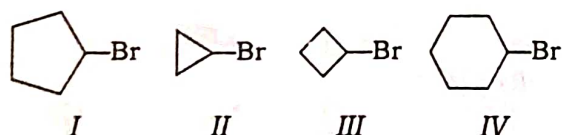
120. The conductivity of 0.001 M dichloroacetic acid solution is $3.88 \times 10^{-5} \text{ S cm}^{-1}$. The molar conductivities at infinite dilution for HCl, NaCl and $\text{CHCl}_2\text{COONa}$ (sodium dichloroacetate) are $426 \text{ S cm}^2 \text{ mol}^{-1}$, $126 \text{ S cm}^2 \text{ mol}^{-1}$ and $88 \text{ S cm}^2 \text{ mol}^{-1}$, respectively. The degree of dissociation and dissociation constant of dichloroacetic acid will be respectively

- (a) 0.2 and 1.1×10^{-6}
- (b) 0.1 and 1.1×10^{-6}
- (c) 0.2 and 1.1×10^{-4}
- (d) 0.1 and 1.1×10^{-5}

79. The major products [X] and [Y] formed in the following reactions are



80. The relative $\text{S}_{\text{N}}2$ reactivity order of the following substrates is



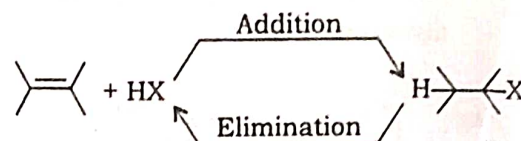
(a) $\text{II} < \text{III} < \text{IV} < \text{I}$

(b) $\text{II} < \text{III} < \text{I} < \text{IV}$

(c) $\text{I} < \text{III} < \text{II} < \text{IV}$

(d) $\text{IV} < \text{I} < \text{III} < \text{II}$

81. Which of the following statements are correct about the following reactions?

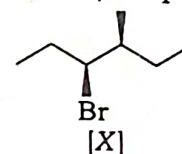


1. Addition is favoured at low temperature whereas elimination is favoured at high temperature.
2. Addition is favoured at high temperature whereas elimination is favoured at low temperature.
3. At high temperature, ΔG is positive for addition reaction whereas ΔG is negative for elimination reaction.
4. At high temperature, ΔG is negative for addition reaction whereas ΔG is positive for elimination reaction.

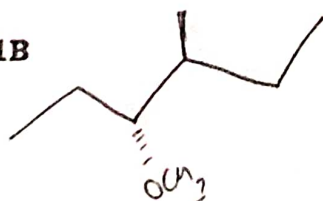
Select the correct answer using the code given below.

- (a) 1 and 3 $\Delta G = \Delta H - T \Delta S$
 (b) 1 and 4
 (c) 2 and 3
 (d) 2 and 4

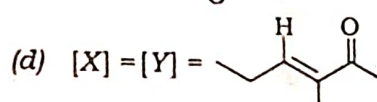
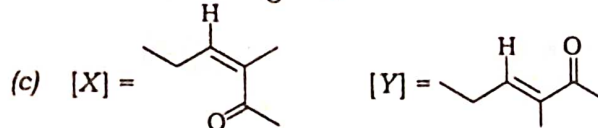
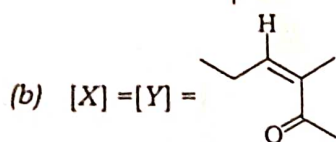
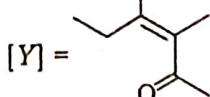
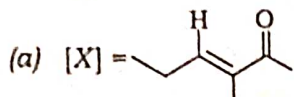
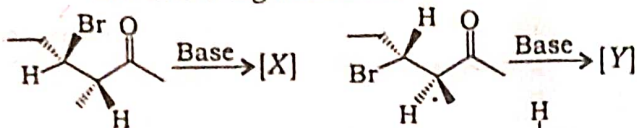
82. The major products formed in $\text{E}2$ and $\text{S}_{\text{N}}2$ reactions of the substrate [X] with sodium methoxide, respectively, are



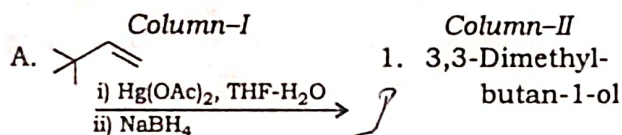
- (a) (Z)-3-methylhex-3-ene and (3R,4R)-3-methoxy-4-methylhexane
- (b) (Z)-3-methylhex-3-ene and (3R,4S)-3-methoxy-4-methylhexane
- (c) (E)-3-methylhex-3-ene and (3R,4S)-3-methoxy-4-methylhexane
- (d) (E)-3-methylhex-3-ene and (3R,4R)-3-methoxy-4-methylhexane



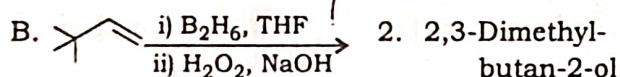
87. The major products [X] and [Y] formed in the following reactions are



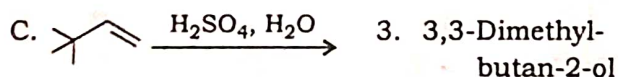
88. Match the reactions from Column-I with the major products formed from Column-II and select the correct answer using the code given below the Columns :



Column-II
1. 3,3-Dimethylbutan-1-ol



2. 2,3-Dimethylbutan-2-ol



3. 3,3-Dimethylbutan-2-ol

4. 2,3-Dimethylbutan-1-ol

Code :

(a)	A	B	C
	2	4	1
(b)	A	B	C
	4	3	2
(c)	A	B	C
	3	1	2
(d)	A	B	C
	3	4	1

89. Consider the following statements :

Statement-1 :

The peroxide effect is observed for the addition of HBr to alkene but not for the addition of HCl or HI.

Statement-2 :

Either of the two propagation steps for the addition of HCl or HI to alkene is endothermic in nature.

Which one of the following is correct with respect to the above statements?

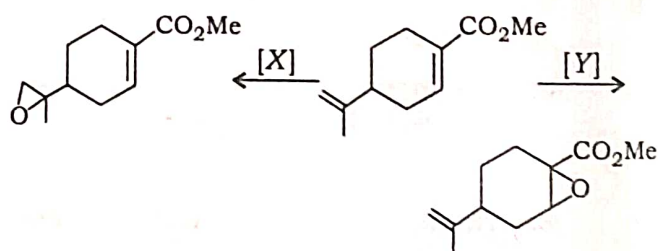
(a) Both Statement-1 and Statement-2 are correct and Statement-2 is the correct explanation of Statement-1.

(b) Both Statement-1 and Statement-2 are correct and Statement-2 is not the correct explanation of Statement-1.

(c) Statement-1 is correct but Statement-2 is incorrect.

(d) Statement-1 is incorrect but Statement-2 is correct.

90. The reagents [X] and [Y] required to carry out the following transformations are



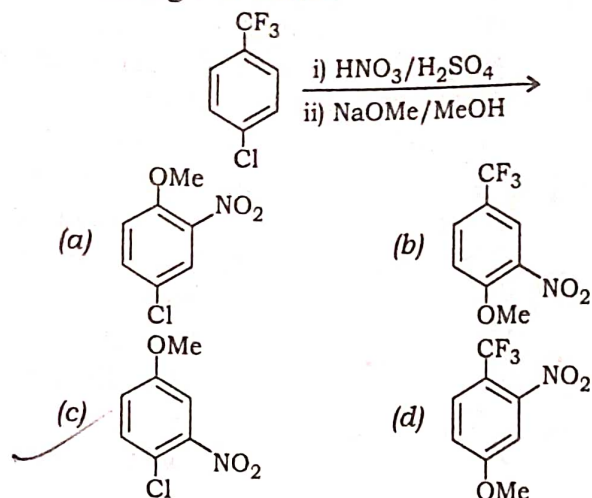
(a) [X] = m-CPBA and [Y] = CH₃CO₃H

(b) [X] = CH₃CO₃H and [Y] = m-CPBA

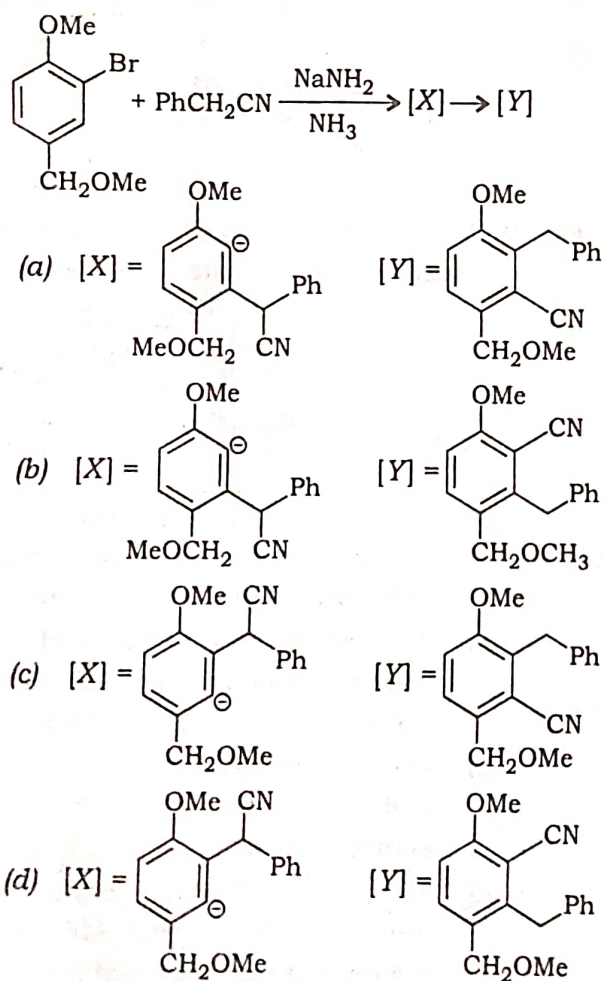
(c) [X] = H₂O₂ / NaOH and [Y] = CH₃CO₃H

(d) [X] = CH₃CO₃H and [Y] = H₂O₂ / NaOH

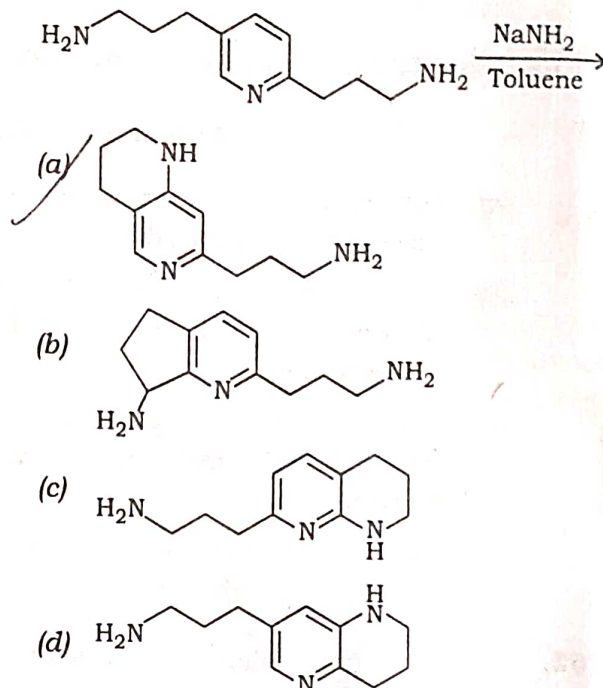
65. The major product formed in the following reaction is



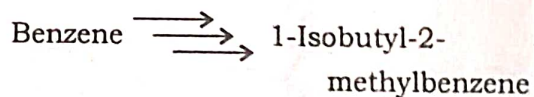
66. The intermediate [X] and the major product [Y] formed in the following reaction are



67. The major product formed in the following reaction is



68. The correct sequence of reactions to accomplish the multi-step synthesis of 1-isobutyl-2-methylbenzene from benzene is



- (a) [1] (CH3)2CHCOCl / AlCl3; [2] HCl, Zn(Hg)/heat; [3] dilute H2SO4; [4] CH3Cl, AlCl3; [5] fuming H2SO4
- (b) [1] CH3Cl, AlCl3; [2] dilute H2SO4; [3] (CH3)2CHCOCl / AlCl3; [4] HCl, Zn(Hg)/heat; [5] fuming H2SO4
- (c) [1] (CH3)2CHCl / AlCl3; [2] fuming H2SO4; [3] CH3Cl, AlCl3; [4] dilute H2SO4
- (d) [1] (CH3)2CHCOCl / AlCl3; [2] HCl, Zn(Hg)/heat; [3] fuming H2SO4; [4] CH3Cl, AlCl3; [5] dilute H2SO4

